

Erratum

A Kinetic and Equilibrium Model for Nylon 6,6 Polymerization

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Lines 24 and 25 on page 2335 read:

where $\Delta H_{\text{app}} = \Delta \overline{H}_L + \Delta \overline{H}_W - \Delta \overline{H}_A - \Delta \overline{H}_C$, $E_{\text{app}} = -(\Delta \overline{H}_A + \Delta \overline{H}_C)$ and T_0 is an arbitrary reference temperature.

They should read:

where $\Delta H_{\text{app}} = \Delta H + \Delta \overline{H}_L + \Delta \overline{H}_W - \Delta \overline{H}_A - \Delta \overline{H}_C$, $E_{\text{app}} = E_a - (\Delta \overline{H}_A + \Delta \overline{H}_C)$, T_0 is an arbitrary reference temperature and ΔH and E_a are the true thermodynamic heat of reaction and activation energy.

Lines 6-8 on page 2343 read:

If one is interested in very high molecular weight nylon 6,6 our results 10% mol fraction water and lower temperatures will yield the most fruitful results.

They should read:

Although the apparent equilibrium constant reaches a maximum at 10% mol fraction water, highest molecular weights at any given temperature are always obtained as the concentration of water goes to zero.

On page 2344 reference 22 reads:

22. K. G. Wyness, *J. Chem. Soc.*, 1958, 2934.

It should read:

22. K. G. Wyness, *J. Chem. Soc.*, 2934 (1958).